

REPLACED BY  
ART 34 AMDT

1

**A PROCESS AND AN APPARATUS FOR PRODUCING CALCIUM  
CARBONATE VIA AN ENZYMATIC PATHWAY****FIELD OF THE INVENTION**

The present invention generally relates to the field of calcium carbonate production  
5 obtained by precipitation. More particularly, it concerns a process for preparing  
precipitated calcium carbonate (hereinafter referred to as PCC) involving the use of  
an enzyme that directly supplies a source of bicarbonate ions which are required  
by the PCC production process. The present invention is particularly useful in the  
pulp and paper industry.

**BACKGROUND OF THE INVENTION**Production of CaCO<sub>3</sub>

Calcium carbonate is used in different fields, most importantly in the pulp and  
paper industry, where it is largely used for coating the paper after it has been  
pressed, thereby increasing the printing quality of the paper, as well as its color,  
15 smoothness, and strength.

The traditional process of forming solid CaCO<sub>3</sub> consists in bubbling carbon dioxide  
in milk of lime (Ca(OH)<sub>2</sub>). This process allows a sufficient amount of CO<sub>2</sub> to  
solubilize and produce solid calcium carbonate according to the following equation:

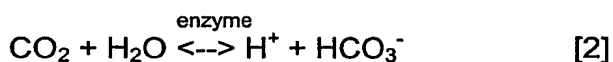


20 This reaction is a rather slow reaction and one drawback thus encountered with the  
same is its low production yield as compared to the need of CaCO<sub>3</sub> in the pulp and  
paper industry or in other fields.

There is thus presently a need for a process for the production of CaCO<sub>3</sub> that  
provides a better production yield.

Conversion of CO<sub>2</sub> into hydrogen ions and bicarbonate ions

EP0991462; AU7753398; WO9855210; CA2291785 in the name of the Applicant disclose the use of the enzyme carbonic anhydrase to catalyse the hydration of CO<sub>2</sub> into hydrogen ions and bicarbonate ions. The reaction that allows for the production of bicarbonate in the presence of the enzyme is represented by equation [2]:



The hydration kinetics of CO<sub>2</sub>, with or without enzyme, has been the object of several scientific works. The reaction rate constant of non-catalysed (without enzyme) reactions is in the order of 0,035s<sup>-1</sup> (25°C, pH=7)<sup>1</sup>. At an elevated pH (pH > 10), the hydroxyl ions (OH<sup>-</sup>) contribute to the catalysis, and consequently, the reaction rate constant can attain 8.5 x 10<sup>3</sup> s<sup>-1</sup> at 25°C<sup>2</sup>. The catalysed reaction (with enzyme), can also be associated with an elevated reaction rate constant in the order of 1,4 x 10<sup>6</sup> (25°C, pH= 8.8)<sup>3</sup>. The ratio of the two reaction rate constants allows one to predict the CO<sub>2</sub> hydration kinetics, which can be in the order of 50 million times more elevated when an enzyme is present and depending on the experimental conditions.

**SUMMARY OF THE INVENTION**

An object of the present invention is to provide a process and an apparatus for the production of CaCO<sub>3</sub> that satisfies the above-mentioned need for a process that provides an increased production yield of precipitated calcium carbonate (PCC).

In accordance with the present invention, that object is achieved with a process for producing CaCO<sub>3</sub>, characterized in that it comprises the steps of:

a) catalysing the hydration of CO<sub>2</sub> contained in a CO<sub>2</sub>-containing gas by means of an enzyme capable of catalysing the hydration of dissolved CO<sub>2</sub> into hydrogen ions and bicarbonate ions, thereby producing a solution containing bicarbonate ions and hydrogen ions;

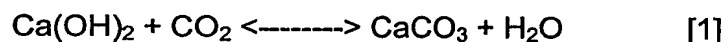
b) reacting the bicarbonate ions contained in the solution obtained in step a) with calcium ions, thereby producing a solution containing  $\text{CaCO}_3$ ; and

c) precipitating the  $\text{CaCO}_3$  contained in the solution obtained in step b).

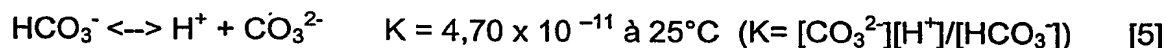
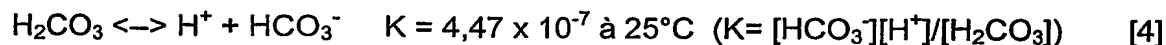
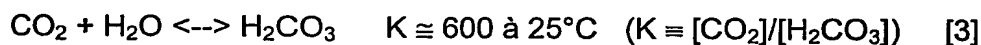
The enzyme used in the process is preferably carbonic anhydrase or an analogue thereof.

In accordance with a preferred aspect, step b) comprises the step of mixing the solution obtained in step a) with  $\text{Ca(OH)}_2$ , thereby providing the calcium ions. In that case, the  $\text{CaCO}_3$  is thus produced from milk of lime ( $\text{Ca(OH)}_2$ ) and  $\text{CO}_2$  as in the prior art, although it shows an improved yield, as explained hereinafter.

As mentioned above, the traditional process of forming  $\text{CaCO}_3$  consists in bubbling carbon dioxide in milk of lime ( $\text{Ca(OH)}_2$ ). This process produces solid calcium carbonate according to the following equation:



This non catalyzed solubilization of gaseous  $\text{CO}_2$  in an aqueous medium implies the following series of natural reversible reactions. The equilibrium constants of these natural reactions are well known in the prior art.

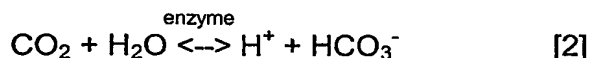


From a thermodynamic point of view, the large equilibrium constant associated to equation [3] translates into a weak inherent tendency to produce the acid ( $K_{\text{eq}} = [\text{CO}_2]/[\text{H}_2\text{CO}_3] \cong 600$ ) (Chemistry of the Elements (2<sup>nd</sup> edition, Butterworth Heinemann, p. 310 (1997))).

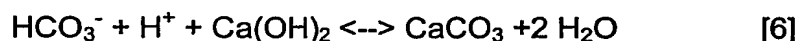
Equation [3], representing the formation of carbonic acid ( $\text{H}_2\text{CO}_3$ ), is in reality the limiting step of the process for producing bicarbonate.

**REPLACED BY  
ART 34 AMDT**

Thus, the use of an enzyme, preferably carbonic anhydrase, which is specific to CO<sub>2</sub> molecules, avoids the step of carbonic acid formation. In other words, the enzyme catalyses the hydration of CO<sub>2</sub> by eliminating the limiting step of the process. As mentioned above, the reaction that allows for the production of bicarbonate in the presence of the enzyme is represented by equation [2]:



Reaction [2] allows for the production of an aqueous solution having a high HCO<sub>3</sub><sup>-</sup> ion content that is used in the process of producing PPC. This concept is represented in the following equation:



From equations [1] and [6], one can see that the reaction for producing PCC performs better in the presence of an enzymatic bioreactor supplied with gaseous CO<sub>2</sub> which in turn directly supplies HCO<sub>3</sub><sup>-</sup> to the process for preparing precipitated calcium carbonate. The improvement resides not only in the increased hydration rate of CO<sub>2</sub>, but also in the increased amounts of CaCO<sub>3</sub> formed in a given reaction time. The variation of Gibbs free energy (ΔG) for each of the reactions will indeed allow one to determine which of the two chemical reactions is favourable from a thermodynamic point of view.

Reaction [6], having a ΔG in the order of -119 kJ/mol, is more favourable from a thermodynamic point of view than reaction [1] which has a ΔG value in the order of -74 kJ/mol. The respective equilibrium constant of these two chemical reactions is calculated using equation [7]:

$$-\Delta G = RT \ln(K_{\text{eq}}) \quad [7]$$

The equilibrium constant of a chemical reaction not containing enzyme is about 9,36 x 10<sup>12</sup> at 298K, while a catalysed reaction making use of an enzyme is about 7,24 x 10<sup>20</sup>. While the equilibrium constant is higher for a reaction containing an

enzyme, the chemical equilibrium is more favourable towards the product ( $\text{CaCO}_3$ ) than to the reagents, and consequently, a better yield is reached.

In accordance with a preferred aspect of the invention, the step a) of  $\text{CO}_2$  hydration is performed in a bioreactor comprising a reaction chamber filled with the enzyme and step b) is performed in at least one separate reaction tank, the process further comprising a step of directing a flow of the solution from the bioreactor into the reaction tank.

Also preferably, step c) of precipitating is performed in the reaction tank. In the present context, precipitating the  $\text{CaCO}_3$  means the formation of separable solid  $\text{CaCO}_3$  from the solution or crystallization of the  $\text{CaCO}_3$ .

### BRIEF DESCRIPTION OF THE FIGURES

Figure 1 is a schematic representation of a first installation suitable for producing precipitated calcium carbonate according to the process of the present invention.

Figure 2 is a schematic representation of a second installation suitable for producing precipitated calcium carbonate according to the process of the present invention.

### DESCRIPTION OF PREFERRED EMBODIMENTS OF THE INVENTION

The invention consists in replacing the step of bubbling gaseous  $\text{CO}_2$  in the milk of lime by a direct input of  $\text{HCO}_3^-$  in an aqueous solution so as to accelerate the production rate of  $\text{CaCO}_3$ .

Referring to figures 1 or 2, a  $\text{CO}_2$  conversion bioreactor (1) having a reaction chamber (2) filled with carbonic anhydrase is used. Preferably, the reaction chamber (2) is filled with packing (3) on which carbonic anhydrase is immobilized.

The most common form which the  $\text{CO}_2$  conversion bioreactor (1) can take is the one containing a packed column where gaseous  $\text{CO}_2$  or other effluents containing  $\text{CO}_2$  penetrate towards the bottom of column (5) at a high concentration ( $> 10\,000$

ppm) and exit by the outlet (6) at a concentration in the order of 500 ppm. As for example, a bioreactor as the one described in WO9855210; CA2291785 is suitable for performing the process. The dimensions of the column depend on both the ratio of gas vs. liquid flow rate and the difference of concentrations between the inlet (5) and the outlet (6). The liquid phase, containing water and a buffer solution so as to allow the pH of the enzymatic medium to be maintained, flows into the bioreactor by an upper inlet (7) and crosses the gas flow while flowing towards the bottom of the reaction chamber where the outlet (8) is located. During the crossing, water, CO<sub>2</sub>, and the enzyme react according to equation [2]. The concentration of gaseous CO<sub>2</sub> then decreases since water fills with HCO<sub>3</sub><sup>-</sup> and H<sup>+</sup> ions in equivalent proportions upon transformation of the CO<sub>2</sub> into HCO<sub>3</sub><sup>-</sup> and H<sup>+</sup> ions. The gas concentration at the outlet (6) is an indicator of the amount of HCO<sub>3</sub><sup>-</sup> dissolved. However, it is preferable to measure the amount of dissolved ions in the solvent (8) found in the pipes.

The solution containing the ions is then directed towards one or more reaction tanks (9 or 9') that contain the milk of lime. It is worth mentioning that the milk of lime is preferably kept at an optimal concentration so as to produce homogenous PCC crystals. Turning now to figure 2, the reaction tanks (9 and 9') are used in parallel and can be numerous. The temperature at which milk of lime is kept can be adjusted in the conditioning tank (10) by means of heating/cooling system (11). This system allows the solution to be either heated or cooled depending on the HCO<sub>3</sub><sup>-</sup> solution properties. Afterwards, an exact amount of milk of lime is added into the reaction tanks (9 and 9') by means of a volumetric pump (4). Indeed, when the reagents are placed together, CaCO<sub>3</sub> crystals are formed within the first moments of contact. The content of the reaction tank (9 and 9') is preferably stirred by means of conventional mixer (12 or 12') to prevent the CaCO<sub>3</sub> crystals from settling.

Finally, the content of each of the reaction tanks (9 and 9') is alternatively transferred one after another into a buffer tank (13) where the liquid and the calcium carbonate solution await to enter the filtration system (15) where they are

to be separated from one another. A mixing system (14) can be used to avoid the formation of deposits (sedimentation) along the bottom of the buffer tank (13). The use of a buffer tank (13) enables the reaction tanks (9 and 9') to remain available to receive solutions so as to keep the process continuous.

- 5 It is worth mentioning that it is also possible to directly add the enzyme in free form into the reaction chamber (9 or 9'). However, in doing so, it can compromise the precipitate purity as well as consuming large amounts of enzyme. Indeed, immobilization of anhydrase on the packing (3) allows the enzyme to be reused many times over until its performance is no longer optimal.
- 10 The process and the apparatus according to the invention are advantageously applicable in the industry of pulp and paper for improving the production yield of  $\text{CaCO}_3$  used for coating the paper.

It is however worth mentioning that the invention could also be advantageous in other fields requiring the use of  $\text{CaCO}_3$ .

- 15 Although preferred embodiments of the present invention have been described in detail herein and illustrated in the accompanying drawings, it is to be understood that the invention is not limited to these precise embodiments and that various changes and modifications may be effected therein without departing from the scope or spirit of the present invention.

20

---

1 (*Carbonic Anhydrase: Zinc and the Mechanisms of Catalysis*. Biology and Chemistry of the Carbonic Anhydrases; Annals of the New-York Academy of Sciences, **429** (1984) pp. 26 to 48, (Coleman J.E.); "The Carbonic Dioxide Hydration activity of carbonic anhydrase I. Stop-Flow Kinetic studies on the Native Human Isoenzymes B and C", *Journal of Biological Chemistry*, **246**(8) (1971), pp. 2561 to 2573 (Khalifah, R.G.); *Biochimie* (2000) DeBoeck Université (Garett *et al.*)

25

2 (*Carbonic Anhydrase: Zinc and the Mechanisms of Catalysis*. Biology and Chemistry of the Carbonic Anhydrases; Annals of the New-York Academy of Sciences, **429** (1984) pp. 26 to 48, (Coleman J.E.); Carbonic Anhydrase Kinetics and Molecular Function: The carbonic Anhydrase, Plenum Press, pp. 49 to 69 (1991), Khalifah *et al.*)

30

3 ("The Carbonic Dioxide Hydration activity of carbonic anhydrase I. Stop-Flow Kinetic studies on the Native Human Isoenzymes B and C", *Journal of Biological Chemistry*, **246**(8) (1971), pp. 2561 to 2573 (Khalifah, R.G.)